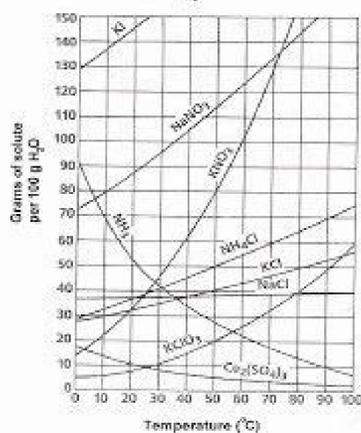


Continue

Answer the following questions based on the solubility curve below.

- Which salt is least soluble in water at 20 °C? KClO₃
- How many grams of potassium chloride can be dissolved in 200 g of water at 80 °C? 100g
- At 40 °C, how much potassium nitrate can be dissolved in 300 g of water? 100g in 200g H₂O → 100g in 300g H₂O
- Which salt shows the least change in solubility from 0 °C to 100 °C? NaCl
- At 30 °C, 80 g of sodium nitrate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? unsaturated
- A saturated solution of potassium chloride is formed from 100 g of water. If the saturated solution is cooled from 80 °C to 50 °C, how many g of precipitate are formed? 20g = 100g - 80g
- What compound shows a decrease in solubility from 0 °C to 100 °C? NH₃, Ca₂(SO₄)₂
- Which salt is most soluble at 10 °C? KI
- Which salt is least soluble at 50 °C? KClO₃
- Which salt is least soluble at 90 °C? Ca₂(SO₄)₂



Name: _____ Date: _____ Class: _____

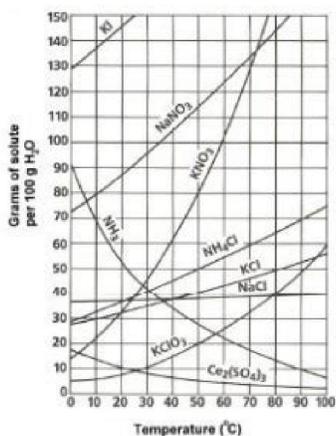
Solubility Curve Practice Problems Worksheet 1

Here's an example of how to read the graph. Find the curve for KClO₃.

At 30°C approximately 10g of KClO₃ will dissolve in 100g of water. If the temperature is increased to 80°C, approximately _____ of the substance will dissolve in 100g (or 100mL) of water.

Directions: Use the graph to answer the following questions. REMEMBER UNITS!

- What is the maximum mass of solute that will dissolve in 100mL of water at the following temperatures?
 - KNO₃ at 70°C = _____
 - NaCl at 100°C = _____
 - NH₄Cl at 90°C = _____
- Which of the above three substances is most soluble in water at 15°C = _____



CHEMISTRY WORKSHEET

Balancing Equations (Part 2)

Name: _____

Write the correctly balanced equation and the type of reaction:

- When strongly heated, magnesium sulfide formed a crystalline deposit on the walls of a sealed reaction vessel. Metallic deposits were found on the bottom of the container.
 Equation: _____
 Type of reaction: _____
- Nitrogen gas reacts with hydrogen gas to form gaseous ammonia.
 Equation: _____
 Type of reaction: _____
- Oxygen gas reacts with solid copper metal to form copper (II) oxide solid.
 Equation: _____
 Type of reaction: _____
- Aluminum metal reacts with copper (II) chloride solution and solid copper.
 Equation: _____
 Type of reaction: _____
- Dinitrogen pentoxide gas forms trinitrogen pentoxide gas and oxygen gas.
 Equation: _____
 Type of reaction: _____
- Hydrogen sulfide gas is bubbled through a solution of iron (III) chloride.
 Equation: _____
 Type of reaction: _____
- Magnesium nitride solid is heated until it forms solid magnesium metal and nitrogen gas.
 Equation: _____
 Type of reaction: _____
- Aluminum solid reacts with oxygen gas to form aluminum oxide solid.
 Equation: _____
 Type of reaction: _____
- Gaseous ammonia is formed from elements.
 Equation: _____
 Type of reaction: _____
- Hydrogen gas and iron (III) oxide powder react to form liquid water and solid iron powder.
 Equation: _____
 Type of reaction: _____

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